ICE Table Practice Problem #2

In the following reaction, $K_{eq} = 9.3 \times 10^{-7}$ at room temp. Calculate the equilibrium concentration of N2O4 in a flask initially containing only 3.00 M of NO_2

$$2\ NO_{2(g)} \rightarrow N_2O_{4(g)}$$

Rxn	$2\ NO_{2(g)}$ →	$N_2O_{4(g)}$
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5%		
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